What does the symbol Δh stand for

VerifiedHint: We must know the difference between atoms, molecules and elements. Then only we can find the name of H.Complete step by step solution:Let’s first discuss all the terms that are given in this question. One atom means that a single atom without any bonds and only having its internal interaction between nuclei and electrons. One molecule means two or more atoms joined together by a bond to form a bigger, stable unit. In this many different types of bonds can be present like ionic bonds (for example\[NaCl\]) or covalent bonds (for example\[C{H_4}\]). Molecules have different characteristics and properties as compared with atoms.One ion means a charged single molecule or atom. An ion can contain either \[+ ve\]or \[-ve\]charge. The charge is acquired due to electron donation or acceptance. If an electron is denoted a \[+ ve\] charge is being acquired and if an electron is being accepted then a \[-ve\] charge is being acquired.Two atoms mean two atoms which are not bonded together or by any other atoms. They are freely moving without any or very little interaction with each other.Now, looking at the question we can see that H cannot be one molecule as a molecule needs two or more than two atoms joined together. Similarly since it doesn’t carry any charge it cannot be an ion. Two atoms is also not possible as no such indication is given. So, the answer will be option A. one atom.Note: We must know that hydrogen is the first element in the periodic table with only one proton and one electron and no neutron. It has an atomic number of 1. It has 2 isotopes namely deuterium and tritium having one and two neutrons respectively. The atomic weight of all three is 1, 2 and 3 respectively.Read LessBook your Free Demo session Photo Courtesy: Kateryna Kon/Science Photo Library/Getty Images The human body is made up of several critical systems — one of which is the digestive system — that are essential to our health. When the digestive system isn’t working properly, the consequences can become uncomfortable and painful very quickly. In some cases, digestive issues go far beyond discomfort and lead to dangerous health conditions. Bacteria like Helicobacter pylori (H. pylori) are often behind gastrointestinal conditions, such as peptic ulcers and even stomach cancer. H. pylori infects the stomach and is particularly dangerous because most people don’t realize they have the bacterial infection until they develop painful symptoms. On the positive side, it’s possible to treat the infection, even years after it starts, but to do so, you need to recognize the potential symptoms of H. pylori.Causes of H. PyloriH. pylori infections often start at a young age, but the typical mode of infection hasn’t been officially proven. Most medical experts believe the bacteria spreads from kissing and other close person-to-person contact or from consuming contaminated food or water. The bacteria are widespread, with large numbers of people all over the world infected with strains. Approximately 50% of people coming to the U.S. from developing Latin countries and Eastern European countries are infected, for example. Photo Courtesy: Westend61/Getty Images The H. pylori bacteria doesn’t discriminate, but those living in certain conditions are more likely to be exposed. People who live with many other people are more likely to contract it, particularly if they live in close quarters. Those in developing countries are also more susceptible, especially if consistent access to clean running water is a problem. Additionally, anyone living with or having close contact with someone else who is infected is almost sure to become infected.Symptoms of H. PyloriSome people appear to be naturally resistant to H. pylori and never experience any symptoms or don’t experience them for many years. This makes it almost impossible to treat H. pylori early before some damage has been done to the stomach and intestines. In most cases, the first indicators of H. pylori infections appear when people start to experience the early symptoms of ulcers or gastritis. Photo Courtesy: Michele Constantini/PhotoAlto Agency RF Collections/Getty Images Gastritis causes redness and swelling in the stomach lining, and ulcers cause actual sores, bleeding and eventually holes in the lining. The milder symptoms of ulcers include bloating, burping, nausea, vomiting, loss of appetite and weight loss. As the ulcers worsen, the symptoms escalate to include anemia and ongoing dull pain in the stomach, especially several hours after eating or when the stomach is empty. Eating or taking antacids typically alleviates the pain for a short time.Diagnosis of H. PyloriWhen patients are diagnosed with ulcers or gastritis, doctors will try to determine the cause, and that includes testing for H. pylori bacteria. Stool sample testing provides information about many different types of bacteria in the intestinal tract, while stool antigen testing looks specifically for the H. pylori bacteria. Certain blood tests can also detect H. pylori antibodies in the blood. Photo Courtesy: Sebastian Kaulitzki/Science Photo Library/Getty Images The H. pylori bacteria makes an enzyme known as urease. This enzyme reduces the acidity of stomach acid and weakens the stomach lining, making it easier for ulcers to develop. Breath tests can look for carbon in the breath, which indicates the presence of urease. In more extreme cases, doctors may perform an upper endoscopy, which involves running a tube with a small camera down the esophagus and into the stomach and upper intestine (duodenum). This allows them to take tissue samples as well as examine potential damage.Treatment of H. PyloriFortunately, H. pylori can be treated at any stage, although the form of treatment varies based on the severity of the condition, the patient’s age and the exact symptoms. It’s common to take multiple medications to ensure the infection is eliminated. Two different antibiotics taken simultaneously can help prevent antibiotic resistance in the bacteria, while acid-reducing medications give the stomach lining a chance to heal. Photo Courtesy: Witthaya Prasongsin/Moment/Getty Images Common acid reducers include histamine (H-2) blockers, proton pump inhibitors and stomach lining protectors. The most common stomach lining protector is Pepto-Bismol, a bismuth subsalicylate, which coats the stomach lining to protect it from acid. H-2 blockers prevent the production of histamine, a substance that helps the stomach make acid. Proton pump inhibitors actually stop the stomach’s acid pump from working, which inhibits the production of acid.Complications of H. PyloriGastritis and peptic ulcers are the most common conditions caused by H. pylori bacteria, although not everyone infected with the bacteria will become ill. Besides pain and other uncomfortable symptoms, ulcers can cause bleeding and holes in the lining of the stomach. Depending on the position of the ulcer, it could also cause stomach blockages. In the worst cases, these bacterial infections can cause gastric cancer, which is the second leading cause of cancer deaths in the world.Photo Courtesy: Burak Karademir/Moment/Getty Images MORE FROM REFERENCE.COM Symbol Meaning + used to separate one reactant or product from another used to separate the reactants from the products - it is pronounced "yields" or "produces" when the equation is read used when the reaction can proceed in both directions - this is called an equilibrium arrow and will be used later in the course (g) indicates that the substance is in a gaseous state an alternative way of representing a substance in a gaseous state (s) indicates that the substance is in a solid state an alternative way of representing a substance in a solid state (aq) indicates that the substance is dissolved in water - the aq comes from aqueous indicates that heat is applied to make the reaction proceed This page explains what an enthalpy change is, and then gives a definition and brief comment for three of the various kinds of enthalpy change that you will come across. Enthalpy change is the name given to the amount of heat evolved or absorbed in a reaction carried out at constant pressure. It is given the symbol ΔH, read as "delta H". Standard enthalpy changes refer to reactions done under standard conditions, and with everything present in their standard states. Standard states are sometimes referred to as "reference states". Standard conditions are: 298 K (25°C) a pressure of 1 bar (100 kPa). where solutions are involved, a concentration of 1 mol dm-3 For a standard enthalpy change everything has to be present in its standard state. That is the physical and chemical state that you would expect to find it in under standard conditions. That means that the standard state for water, for example, is liquid water, H2O(l) - not steam or water vapour or ice. Oxygen's standard state is the gas, O2(g) - not liquid oxygen or oxygen atoms. For elements which have allotropes (two different forms of the element in the same physical state), the standard state is the most energetically stable of the allotropes. For example, carbon exists in the solid state as both diamond and graphite. Graphite is energetically slightly more stable than diamond, and so graphite is taken as the standard state of carbon. Similarly, under standard conditions, oxygen can exist as O2 (simply called oxygen) or as O3 (called ozone - but it is just an allotrope of oxygen). The O2 form is far more energetically stable than O3, so the standard state for oxygen is the common O2(g). The symbol for a standard enthalpy change is ΔH°, read as "delta H standard" or, perhaps more commonly, as "delta H nought". Remember that an enthalpy change is the heat evolved or absorbed when a reaction takes place at constant pressure. The standard enthalpy change of a reaction is the enthalpy change which occurs when equation quantities of materials react under standard conditions, and with everything in its standard state. That needs exploring a bit. Here is a simple reaction between hydrogen and oxygen to make water: First, notice that the symbol for a standard enthalpy change of reaction is ΔH°r. For enthalpy changes of reaction, the "r" (for reaction) is often missed off - it is just assumed. The "kJ mol-1" (kilojoules per mole) doesn't refer to any particular substance in the equation. Instead it refers to the quantities of all the substances given in the equation. In this case, 572 kJ of heat is evolved when 2 moles of hydrogen gas react with 1 mole of oxygen gas to form 2 moles of liquid water. Notice that everything is in its standard state. In particular, the water has to be formed as a liquid. And there is a hidden problem! The figure quoted is for the reaction under standard conditions, but hydrogen and oxygen don't react under standard conditions. Whenever a standard enthalpy change is quoted, standard conditions are assumed. If the reaction has to be done under different conditions, a different enthalpy change would be recorded. That has to be calculated back to what it would be under standard conditions. Fortunately, you don't have to know how to do that at this level. The standard enthalpy change of formation of a compound is the enthalpy change which occurs when one mole of the compound is formed from its elements under standard conditions, and with everything in its standard state. The equation showing the standard enthalpy change of formation for water is: When you are writing one of these equations for enthalpy change of formation, you must end up with 1 mole of the compound. If that needs you to write fractions on the left-hand side of the equation, that is OK. (In fact, it is not just OK, it is essential, because otherwise you will end up with more than 1 mole of compound, or else the equation won't balance!) The equation shows that 286 kJ of heat energy is given out when 1 mole of liquid water is formed from its elements under standard conditions. Standard enthalpy changes of formation can be written for any compound, even if you can't make it directly from the elements. For example, the standard enthalpy change of formation for liquid benzene is +49 kJ mol-1. The equation is: If carbon won't react with hydrogen to make benzene, what is the point of this, and how does anybody know what the enthalpy change is? What the figure of +49 shows is the relative positions of benzene and its elements on an energy diagram: How do we know this if the reaction doesn't happen? It is actually very simple to calculate it from other values which we can measure - for example, from enthalpy changes of combustion (coming up next). We will come back to this again when we look at calculations on another page. Knowing the enthalpy changes of formation of compounds enables you to calculate the enthalpy changes in a whole host of reactions and, again, we will explore that in a bit more detail on another page. And one final comment about enthalpy changes of formation: The standard enthalpy change of formation of an element in its standard state is zero. That's an important fact. The reason is obvious . . . For example, if you "make" one mole of hydrogen gas starting from one mole of hydrogen gas you aren't changing it in any way, so you wouldn't expect any enthalpy change. That is equally true of any other element. The enthalpy change of formation of any element has to be zero because of the way enthalpy change of formation is defined. The standard enthalpy change of combustion of a compound is the enthalpy change which occurs when one mole of the compound is burned completely in oxygen under standard conditions, and with everything in its standard state. The enthalpy change of combustion will always have a negative value, of course, because burning always releases heat. Two examples: Notice: Enthalpy of combustion equations will often contain fractions, because you must start with only 1 mole of whatever you are burning. If you are talking about standard enthalpy changes of combustion, everything must be in its standard state. One important result of this is that any water you write amongst the products must be there as liquid water. Similarly, if you are burning something like ethanol, which is a liquid under standard conditions, you must show it as a liquid in any equation you use. Notice also that the equation and amount of heat evolved in the hydrogen case is exactly the same as you have already come across further up the page. At that time, it was illustrating the enthalpy of formation of water. That can happen in some simple cases. Talking about the enthalpy change of formation of water is exactly the same as talking about the enthalpy change of combustion of hydrogen. Jim Clark (Chemguide.co.uk)

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